

PHYSICS

1. 2
2. 4
3. 4

4. 4 $C = \sin^{-1}\left(\frac{v_1}{v_2}\right) = \sin^{-1}\left(\frac{1.8 \times 10^8}{2.4 \times 10^8}\right) = \sin^{-1}\left(\frac{3}{4}\right)$

5. 1 $n = \frac{\sin(i)}{\sin\left(\frac{i}{2}\right)} = \frac{2 \sin\left(\frac{i}{2}\right) \cos\left(\frac{i}{2}\right)}{\sin\left(\frac{i}{2}\right)}$; $n = 2 \cos\left(\frac{i}{2}\right)$; $i = 2 \cos^{-1}\left(\frac{n}{2}\right)$

6. 3 $\sin c = \frac{n_2}{n_1} = \frac{\lambda_1}{\lambda_2}$; $\sin c = \frac{3500}{7000} = \frac{1}{2} \therefore c = 30^\circ$

7. 3

8. 2 30°

$\tan p = n = \sqrt{3}$; $p = 60^\circ$; At the polarising angle reflected and refracted beams are mutually perpendicular

9. 2 Slope = h/e . Here, h and e are constants. Therefore, the slope is same for all metals and independent of the intensity of incident radiation.

10. 1 We have

$$E_k = hf - hf_0$$

and $(E_k)' = 2hf - hf_0$

Therefore, $\frac{(E_k)' + E_k}{E_k} = \frac{3hf - 2hf_0}{hf - hf_0} = \frac{3f - 2f_0}{f - f_0}$

$$\frac{(E_k)'}{E_k} = \frac{3f - 2f_0}{f - f_0} - 1 = \frac{3f - 2f_0 - f + f_0}{f - f_0}$$

$$\frac{(E_k)'}{E_k} = \frac{2f - f_0}{f - f_0} = \frac{2f - 2f_0}{f - f_0} + \frac{f_0}{f - f_0}$$

$$\frac{(E_k)'}{E_k} = 2 + \left(\frac{f_0}{f - f_0}\right) E_k$$

Therefore, $(E_k)' = 2E_k + \left(\frac{f_0}{f - f_0}\right) E_k$

11. 2

12. 2 p-type is obtained by doping with trivalent impurity atoms

13. 2

14. 3 The intermediate image in a compound microscope is real, inverted and magnified.

15. 1 $M = \frac{f_o}{f_e} = \frac{200}{5} = 40$

16. 1 $F \propto q_1 q_2$

Ist case

IInd case

$q_1 = +7\mu C$

$q_1 = +7\mu C - 2\mu C = +5\mu C$

$q_2 = -5\mu C$

$q_2 = -5\mu C - 2\mu C = -7\mu C$

17. 3 $\tau_{max} = PE = q(2l)E$

18. 2

19. 4

20. 3 $B = \frac{\mu_0 i}{2\pi r}$

21. 3

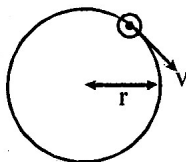
22. 1

Magnetic moment $\mu = niA$

Where n = number of turns of the current loop

i = current; Since the orbiting electron behaves as current loop of current i; we can write

$$i = \frac{e}{T} = \frac{e}{\frac{2\pi r}{v}} = \frac{ev}{2\pi r}$$



A = area of the loop = πr^2

$$\Rightarrow \mu = (1) \left(\frac{ev}{2\pi r} \right) (\pi r^2) \Rightarrow \mu = \frac{evr}{2}$$

23. 4 $x \propto \frac{1}{T}$

24. 1

25. 2

26. 4 We have

$$\lambda = \frac{h}{mv}$$

$$m_\beta < m_p < m_n < m_\alpha \Rightarrow \lambda_\beta < \lambda_p < \lambda_n < \lambda_\alpha$$

27. 2 $r \propto n^2$

28. 2 No. of spectral lines = $\frac{n(n-1)}{2} = \frac{4(4-1)}{2} = 6$

29. 2 B.E. = Δmc^2 and B.E. per nucleons = $\frac{\text{B.E.}}{\text{mass no.}}$

30. 2 $r \propto A^{1/3} \Rightarrow \frac{r_1}{r_2} = \left(\frac{A_1}{A_2}\right)^{1/3}$

$\Rightarrow \frac{3}{5} = \left(\frac{27}{A}\right)^{1/3} \Rightarrow \frac{27}{125} = \frac{27}{A}$

$A = 125$

Number of nuclei in atom X = $A - 52 = 125 - 52 = 73$

CHEMISTRY

31. 3 Only chlorobenzene and bromobenzene are prepared by Sand Meyer's reaction
32. 3
33. 2
34. 1
35. 3 Carbonyl compounds containing α -hydrogen atom give aldol condensation
36. 4 Lactose and maltose are disaccharides
37. 4
38. 3 D and L have no relation with the optical rotation. Carbohydrate having D-configuration may be either dextrorotatory or laevorotatory e.g., D-(+)-glucose is dextrorotatory while D-(-)-fructose is laevorotatory.
39. 3
40. 2
41. 4 Amines have an unshared pair of electrons on nitrogen atom due to which they behave as Lewis bases
42. 4 Aniline itself does not undergo Friedel-Crafts reaction because amino group of aniline acts as a base and forms salt with Lewis acid, AlCl_3 etc, $\text{C}_6\text{H}_5\overset{\oplus}{\text{N}}\text{H}_2 \cdot \text{AlCl}_3^-$

43. 4 We know that, $K_b = \frac{\Delta T_b \times W_1 \times MW_2}{1000 \times W_2}$

For CCl_4 ; $K_b = \frac{0.201 \times 50 \times 128}{1000 \times 0.2563} = 5.019$

$MW_2 = \frac{1000 \times K_b \times W_2}{\Delta T_b \times W_1} = \frac{1000 \times 5.019 \times 0.6216}{0.647 \times 50} = 96.44 \text{ g}$

44. 2 The solution containing *n*-heptane and ethanol shows non-ideal behaviour with positive deviation from Raoult's law. This is because the ethanol molecules are held together by strong H-bonds, however the forces between *n*-heptane and ethanol are not very strong, as a result they easily vapourise showing higher vapour pressure than expected.

45. 4

46. 3

47. 2

48. 4

49. 4

50. 1

51. 1

52. 4

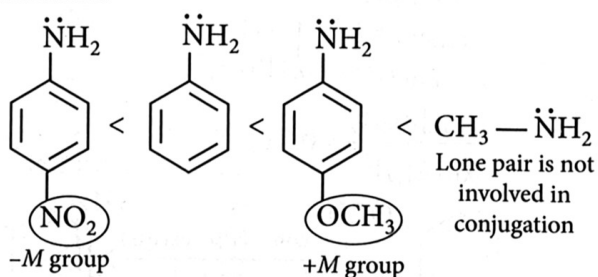
53. 1 In KMnO_4 Mn is in +7 oxidation state

54. 4

55. 3

56. 3

57. 4 Aromatic amines like aniline are less basic than aliphatic amines because of the involvement of lone pair of electrons in resonance with the aromatic ring which now becomes less available for donation. In substituted aromatic amines, electron withdrawing groups decrease the basic character and electron releasing groups increase the basic character.



58. 4

59. 1

60. 2

BIOLOGY

61. 4

62. 2

- 63. 2
- 64. 2
- 65. 3
- 66. 4
- 67. 3
- 68. 1
- 69. 1
- 70. 4
- 71. 2
- 72. 4
- 73. 2
- 74. 3
- 75. 1
- 76. 4
- 77. 2
- 78. 1
- 79. 3
- 80. 4
- 81. 2
- 82. 1
- 83. 2
- 84. 2
- 85. 1
- 86. 4
- 87. 2
- 88. 3
- 89. 1
- 90. 2

MATHEMATICS

- 91. 1
- 92. 1
- 93. 1
- 94. 3
- 95. 2
- 96. 4
- 97. 1
- 98. 2
- 99. 1

100. 4
101. 1
102. 4
103. 4
104. 1
105. 2
106. 3
107. 4
108. 2
109. 3
110. 4
111. 2
112. 1
113. 3
114. 1
115. 1
116. 1
117. 3
118. 4
119. 1
120. 2